

### Key Words

<b>Acid</b>	When dissolved in water, its solution has a pH value less than 7	<b>Neutralisation</b>	The chemical reaction of an acid with a base in which a salt is formed
<b>Alkali</b>	Its solution has a pH value more than 7	<b>Titration</b>	Process to find out how much of a chemical there is in a solution by addition of another liquid of a known strength until an end point is reached
<b>Corrosive</b>	Reacts with materials and makes them dissolve	<b>Base</b>	A substance that neutralises an acid –those that dissolve in water are
<b>pH</b>	Scale of acidity and alkalinity from 0 to 14	<b>Litmus</b>	Indicator solution
<b>Indicator</b>	Substance used to identify whether an unknown solution is acidic or alkaline	<b>irritant</b>	Something that irritates and reddens the skin
<b>neutral</b>	A solution with a pH value of 7; it is neither an acid nor an alkali	<b>concentration</b>	A measure of the number of particles in a given volume.

### Learning Sequence

1. Exploring acids
2. Exploring alkalis
3. Using indicators
4. Exploring neutralisation
5. Investigating neutralisation
6. Investigation part 2

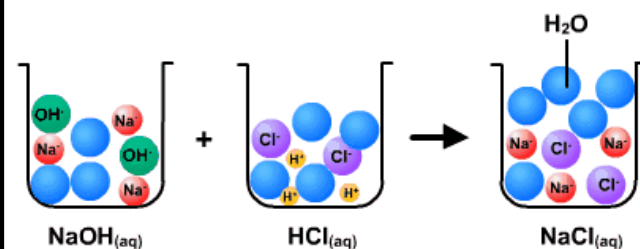
### Assessment

- End of topic test
- Investigation write up

### Indicators

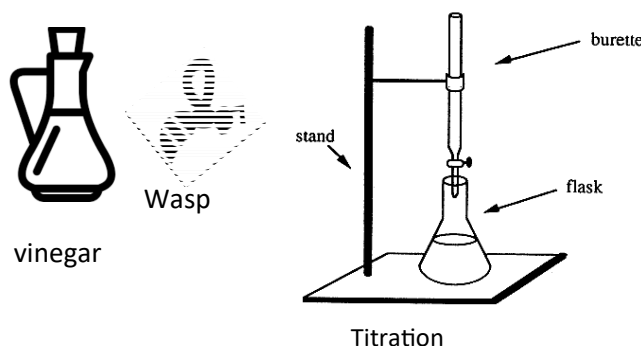
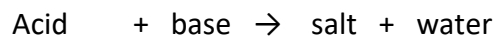


### Neutralisation

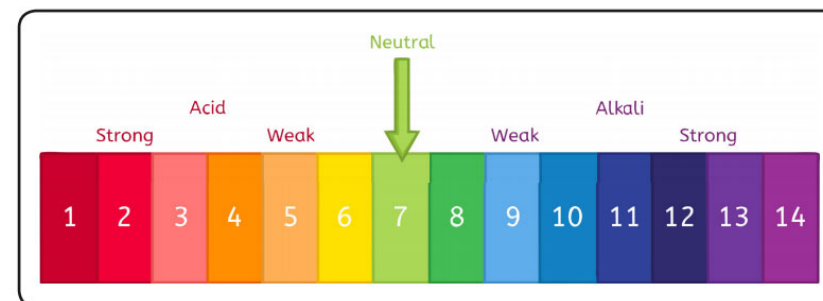


Sodium hydroxide + hydrochloric acid → sodium chloride + water

### General formula:



### Acids, alkalis and pH



hydrogen ions (H <sup>+</sup> )	water (H <sub>2</sub> O)	hydroxide ions (OH <sup>-</sup> )
hydrochloric acid	neutralisation	sodium hydroxide
sulphuric acid	salt	potassium hydroxide
nitric acid		ammonia

### Hazard Symbols

